

Mechanism-Based Inhibition of Myeloperoxidase by Hydrogen Peroxide: Enhancement of Inactivation Rate by Organic Donor Substrates

Femi J. Olorunniji^{1,4,*}, Martin O. Iniaghe², Joseph O. Adebayo¹, Sylvia O. Malomo¹ and Suara A. Adediran^{3,5,*}

¹Department of Biochemistry, Faculty of Science, University of Ilorin, Ilorin, Nigeria

²Department of Biochemistry, Faculty of Science, Ambrose Alli University, Ekpoma, Nigeria

³Department of Chemistry, Faculty of Science, University of Ilorin, Ilorin, Nigeria

⁴Division of Molecular and Cellular Biology, Faculty of Biomedical and Life Sciences, University of Glasgow, Glasgow G12 8QQ, U.K

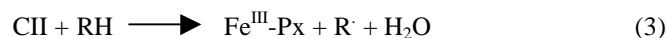
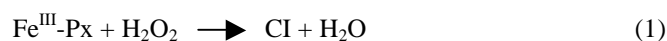
⁵Department of Chemistry, Wesleyan University, Middletown, CT 06459, USA

Abstract: The effects of two hydrazine derivatives (isoniazid and hydralazine) on the inactivation of myeloperoxidase by H₂O₂ were investigated. Incubation of 20 nM myeloperoxidase with 0.25 mM H₂O₂ alone caused a time-dependent irreversible loss of tetramethylbenzidine oxidation activity with a pseudo-first order inactivation rate constant of 0.057 min⁻¹. The hydrazine derivatives increased the inactivation rate in a concentration-dependent manner. Inactivation of the enzyme by H₂O₂ with or without the hydrazides showed a saturation kinetics pattern. Steady state kinetics analysis suggests that the hydrazides likely inactivate myeloperoxidase using a similar inactivating species as does H₂O₂. A bimolecular rate constant, specific inactivation rate enhancement factor (k*_{enh}) is proposed as a formal description of the inactivation rate stimulation by the hydrazides. This parameter potentially avoids confounding the finite inactivation due to H₂O₂ with that caused by the presence of the hydrazides. The relevance of these findings and the constants derived to the analysis of suicide inactivation of peroxidases by reductant substrates are discussed.

Keywords: Myeloperoxidase, Isoniazid, Hydralazine, Mechanism-based inhibition, Inactivation kinetics.

INTRODUCTION

The peroxidases (donor: H₂O₂ oxidoreductase) are haem proteins, which mediate the one-electron oxidation of a wide array of organic compounds (phenols, anilines, β-diketones etc) with the concomitant reduction of H₂O₂ to H₂O [1]. The catalytic cycles of most peroxidases resemble that described for horseradish peroxidase (HRP) [1]. In this scheme, the ferric peroxidase (Fe^{III}-Px) donates two electrons to H₂O₂ resulting in cleavage of H₂O₂ and formation of a redox intermediate of the enzyme called compound I (CI). CI consists of an oxoferryl protein cation radical, in which one of the oxidation equivalents exists as the ferryl ion, Fe (IV) and the other as a porphyrin-centred cation radical. CI reacts with reductants (RH) to generate substrate free radicals and compound II (CII), another redox intermediate in which the oxoferryl species remains intact but the cation radical is reduced. A one-electron reduction of CII by a second RH molecule regenerates the ferric enzyme and forms a second equivalent of R.



Another redox intermediate, Compound III (CIII) is formed in the course of peroxidase catalytic cycle. It is catalytically inactive and exists as a resonance form between the Fe^{II}-O₂ and Fe^{III}-O₂⁻ complexes.

Myeloperoxidase (MPO) is a heterodimeric, cationic and glycosylated haem enzyme. The enzyme is a 140-kDa dimer of identical halves, each consisting of two polypeptide chains of 108 and 466 amino acids. Each half contains a covalently attached haem [2]. X-ray crystallographic studies show that H95 and R239 are likely involved in the catalytic mechanism of MPO [3]. Like the other haem peroxidases, MPO combines with hydrogen peroxide to form the highly reactive redox intermediate called compound I. The high redox potential of this intermediate confers MPO with the unusual ability to oxidize chloride ion to hypochlorous acid with the concomitant regeneration of native enzyme [4].

MPO differs significantly from other peroxidases in its chemical and spectroscopic properties. For instance, its Soret band is red-shifted in the optical absorption and in its pyridine haemochrome spectrum. While there are spectral and kinetics differences between MPO and other mammalian

*Address correspondence to these authors at the Division of Molecular & Cellular Biology, Faculty of Biomedical & Life Sciences, University of Glasgow, Glasgow G12 8QQ, United Kingdom; Tel: +441413303331; E-mail: f.olorunniji@bio.gla.ac.uk and Department of Chemistry, Wesleyan University, Middletown, Connecticut 06459, USA; E-mail: sadediran@wesleyan.edu

peroxidases, the active site amino acid sequence of the enzymes share a high degree of homology. Virtually all known peroxidases are inactivated by H_2O_2 and other hydroperoxides at relatively high concentrations [1]. This substrate inactivation leads to the modification of the haem prosthetic group and formation of a verdohaemoprotein as the final product. The mechanism of the inactivation by hydrogen peroxide is yet to be completely resolved. There appears, however, to be a consensus that compound III formation is crucial to the process [5]. The existence of compound III as the peroxy iron (III) porphyrin free radical resonant form can facilitate the transfer of electrons from the ferrous state to an extra hydrogen peroxide molecule thereby generating hydroxyl radical. This highly reactive species has the propensity to attack the haem porphyrin ring and lead to irreversible inactivation [6].

Some reports suggest that inactivation of MPO by H_2O_2 is not an effective process and that the presence of a radical generating reducing substrate is needed for appreciable inactivation to occur [7-9]. Inactivation of peroxidases by certain reductants is a feature of peroxidase catalysis. Such substrates are oxidized in the presence of H_2O_2 , a process that is concomitant with enzyme inactivation. In this study, we have used a steady state kinetics approach to demonstrate that MPO is irreversibly inactivated by excess H_2O_2 in the absence of reducing substrate. The rate of the inactivation process is enhanced in the presence of suicide substrates, isoniazid (ISD) and hydralazine (HDL), which are known inactivators of the enzyme [10, 11]. We also describe a novel bimolecular rate constant that formally describes the activating effect of a suicide donor substrate on hydroperoxide-initiated inactivation of peroxidases.

MATERIALS AND METHODS

Reagents, Substrates and Enzyme

ISD, HDL, H_2O_2 and horseradish peroxidase were products of Sigma-Aldrich Company (Dorset, Poole, UK). Other chemicals and reagents used were of high quality research grade. MPO was isolated and purified from human polymorphonuclear neutrophils by means of a combination of published procedures with slight modifications [2, 12]. The enzyme preparation had a purity index (A_{428}/A_{280}) of 0.70. The concentration of MPO, expressed as haem concentration was measured spectrophotometrically based on Soret absorbance at 428 nm by using a millimolar absorption coefficient per haem of $89 \text{ mM}^{-1}\text{cm}^{-1}$ [13].

Determination of Enzyme Activity

MPO activity was determined by measuring the rate of oxidation of tetramethylbenzidine (TMB) at 20°C in a 3.0 ml reaction mixture containing 80 mM sodium acetate buffer, pH 5.4, 1.6 mM TMB and 0.3 mM H_2O_2 . The initial rate of formation of the bluish charge transfer complex was measured at 655 nm [14]. Reactions were initiated by adding H_2O_2 as the last component of the mixture. Initial rate of TMB oxidation was determined from the linear portion of the time course for the appearance of the charge transfer complex at 655 nm [15]. Desired concentrations of H_2O_2

were prepared daily from a 30% stock solution and determined spectrophotometrically using horseradish peroxidase as catalyst for I_3^- generation from I^- [16]. Concentrations determined this way were confirmed by absorbance measurements at 240 nm, taking E_{240} as $43.6 \text{ M}^{-1}\text{cm}^{-1}$ [17]. Kinetics absorbance measurements were performed with a UV/visible spectrophotometer model Camspec M105.

Inactivation Kinetics

The kinetics of irreversible inactivation of MPO by H_2O_2 and the hydrazides were studied under pseudo-first order conditions. MPO (20 nM) was incubated with the hydrazides (0.25-1.50 mM) at 20°C in 100 mM phosphate buffer (pH 7.4) in the presence of H_2O_2 (0.25 mM) in a total reaction volume of 1.0 ml. Incubation of MPO with H_2O_2 (0.25 mM) without the hydrazides was carried out under similar conditions. In all cases, H_2O_2 was added last to the incubation mixture containing MPO and hydrazide to initiate the inactivation reaction. At 3 minute intervals, aliquots of 50 μl were withdrawn and diluted 500-fold, and the residual activities (A) were determined. This treatment would prevent the interpretation of 'reversibly inactivated' enzyme in the form of compound III as completely inactivated enzyme [18]. The activity of an equal amount of the enzyme without pre-incubating with H_2O_2 and the hydrazides was determined and taken as the control (A_0). Plots of $\ln(A/A_0)$ against time were used to obtain the pseudo first order inactivation rate constants (k_{obs}). The concentration dependence of H_2O_2 on inactivation rate in the presence and absence of both hydrazides was studied over a H_2O_2 concentration range of 0.1-1.25 mM. MPO (20 nM) was incubated with the appropriate H_2O_2 concentration in 100 mM phosphate buffer, pH 7.4 at 20°C in a 1.0 ml reaction volume. First order inactivation rates (k_{obs}) were determined as described above and the experiments were carried out with and without 0.5 mM ISD or HDL. The H_2O_2 and hydrazide concentrations used in this study were at least ten times that of the enzyme. This was to ensure that pseudo-first order conditions were in operation in the inactivation processes. At least four k_{obs} determinations were made for each H_2O_2 and hydrazide concentrations. Results of variation of k_{obs} with H_2O_2 concentration with or without the hydrazides were fitted to a rectangular hyperbola using non-linear regression generated by a computer programme, Enzfitter (Biosoft Corp.). Double reciprocal transformation of the data was carried out using the same software to obtain Kitz-Wilson inactivation constants, k_{inact} and K_i according to the following equation

$$1/k_{\text{obs}} = K_i/k_{\text{inact}}(1/[\text{H}_2\text{O}_2]) + 1/k_{\text{inact}} \quad (4)$$

Two second order rate constants that describe the effect of the hydrazides on the inactivation by H_2O_2 were proposed in this study. Inactivation rate enhancement factor (k_{enh}) and specific inactivation rate enhancement factor (k_{enh}^*) were calculated from the slopes of the plots of k_{obs} and $k_{\text{hyd}}-k_{\text{per}}$ versus hydrazide concentrations respectively.

RESULTS

We investigated the inactivation of H_2O_2 by MPO in the absence of a reductant substrate by incubating 20 nM of the

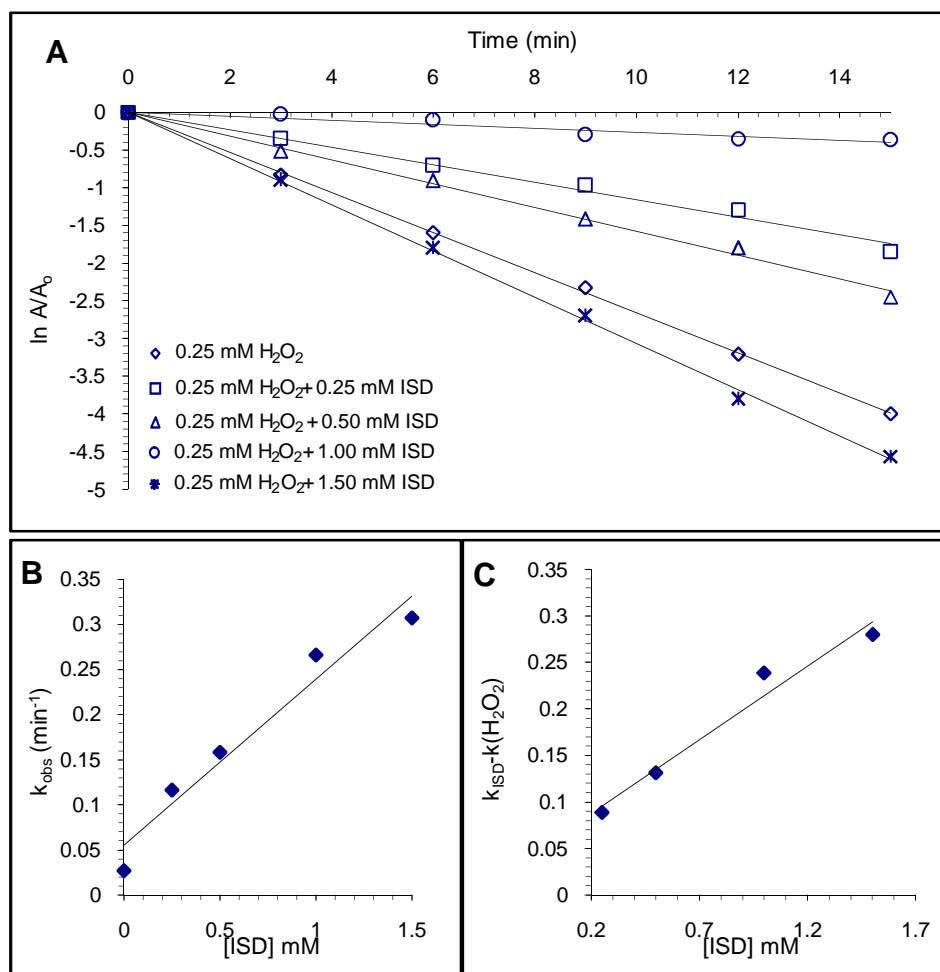


Fig. (1). Rate enhancement of H₂O₂-dependent inactivation of MPO by ISD. Panel A shows the time course of MPO inactivation by 0.25 mM H₂O₂ alone and in combination with 0.25 -1.50 mM ISD. The lines are plots of $\ln A/A_0$ versus time where A_0 and A are activities at time zero and at the corresponding times indicated on the plots respectively. The slopes of the lines represent the first order inactivation rates (k_{obs}) and were plotted versus ISD concentration in Panel B to obtain the inactivation rate enhancement factor (k_{enh}). Panel C shows the determination of the bimolecular inactivation rate enhancement constant (k_{enh}^*) by plotting $k_{\text{ISD}} - k_{\text{per}}$ versus ISD concentration. Reaction conditions are as follows. MPO (20 nM) was incubated with the hydrazides (0.25-1.50 mM) at 20 °C in 100 mM phosphate buffer (pH 7.4) in the presence of H₂O₂ (0.25 mM) in a total reaction volume of 1.0 ml. A similar incubation was carried out with H₂O₂ (0.25 mM) without the hydrazides under similar conditions. In all cases, H₂O₂ was added last to initiate the inactivation reaction. At 3-minute intervals, aliquots of 50 μl were withdrawn and diluted 500-fold before determining the residual activities (A). Activity of an equal amount of the enzyme without pre-incubating with H₂O₂ and the hydrazides was determined and taken as the control (A_0). Plots of $\ln(A/A_0)$ against time were used to obtain the pseudo first order inactivation rate constants (k_{obs}).

enzyme with 0.25 mM hydrogen peroxide and the residual activity was determined at regular intervals. Fig. (1A) shows that 0.25 mM H₂O₂ caused a time-dependent irreversible loss of MPO activity. A pseudo-first order rate constant of 0.116 min⁻¹ was calculated from the plot for the inactivation of MPO by H₂O₂ alone. Inclusion of ISD (0.25-1.50 mM) in the incubation mixture led to increase in the inactivation rate. The inactivation process retained a pseudo-first order kinetics pattern in the presence of ISD. The inactivation rate enhancement by ISD was concentration dependent as shown in Fig. (1B). The time and concentration dependence of ISD

effect on MPO inactivation suggests that the process is mechanism-based and describable in definite kinetics terms. Fig. (2) shows the effect of HDL on MPO inactivation by H₂O₂ as described above for ISD. There is a linear relationship between inactivation enhancement rates and HDL concentration as was observed for ISD. However, HDL is less effective than ISD.

The hydrazides showed concentration-dependent effects on the inactivation of MPO by H₂O₂. Fig. (1B) is a plot of the first order rate inactivation constant against ISD

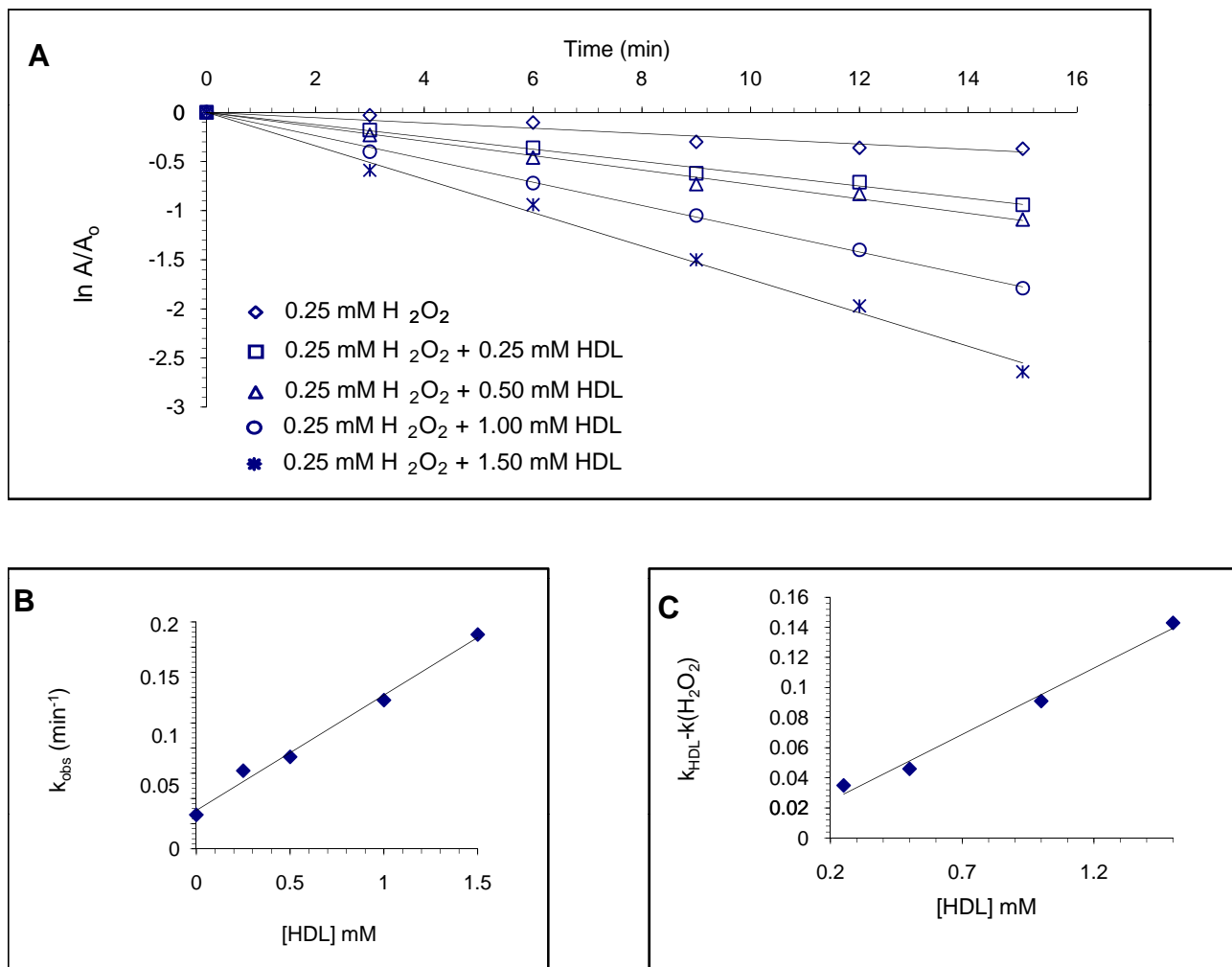


Fig. (2). Rate enhancement of H₂O₂-dependent inactivation of MPO by HDL. Panel A shows the time course of MPO inactivation by 0.25 mM H₂O₂ alone and in combination with 0.25-1.50 mM HDL. The lines are plots of ln A/A₀ versus time where A₀ and A are activities at time zero and at the corresponding time indicated on the plots respectively. The slopes of the lines represent the first order inactivation rates (k_{obs}) and were plotted versus HDL concentration in Panel B to obtain the inactivation rate enhancement factor (k_{enh}). Panel C shows the determination of the bimolecular inactivation rate enhancement constant (k*_{enh}) by plotting k_{hyd}-k_{per} versus HDL concentration. Reaction conditions are as described in Fig. (1).

concentration. There is a linear correlation between k_{obs} and ISD concentration. From the slope of this plot, a *bimolecular inactivation rate enhancement factor* (k_{enh}) of 0.184 min⁻¹ mM⁻¹ was calculated. Fig. (2B) shows a similar plot for the rate enhancement due to HDL with a k_{enh} value of 0.091 min⁻¹ mM⁻¹. This parameter describes in a quantitative manner the effect of the modifiers on the inactivation occasioned by H₂O₂. This linear fit may suggest that the effect of hydrazide on H₂O₂ induced inactivation is a simple chemical event.

In order to separate the inactivation due to H₂O₂ alone from that caused by H₂O₂ in the presence of the hydrazides, another parameter k*_{enh} was expressed. This parameter described here as *specific bimolecular inactivation rate enhancement factor*, was obtained from the slope of the plot of k_{hyd}-k_{per} against the concentrations of the hydrazide, [hyd].

$$k^*_{\text{enh}} = (k_{\text{hyd}} - k_{\text{per}}) / [\text{hyd}] \quad (5)$$

where k_{hyd} and k_{per} are the first order inactivation rate constants in the presence and absence of hydrazide respectively.

The specific inactivation enhancement rates k*_{enh} calculated from the slopes of Figs. (1C and 2C) were 0.159 min⁻¹ mM⁻¹ and 0.088 min⁻¹ mM⁻¹ for ISD and HDL respectively. Going by these calculations, ISD was about 1.7 times more effective than HDL in enhancing the inactivation of MPO by H₂O₂. The validity of k*_{enh} is based on an assumption that the effects of hydrogen peroxide and ISD are mechanistically similar enough to be described by the same microscopic rate constants and that the different effects are additive in nature. Though the validity of this primary assumption is yet to be established, Figs. (1C and 2C) show that there is a linear correlation between k*_{enh} and ISD or HDL concentration.

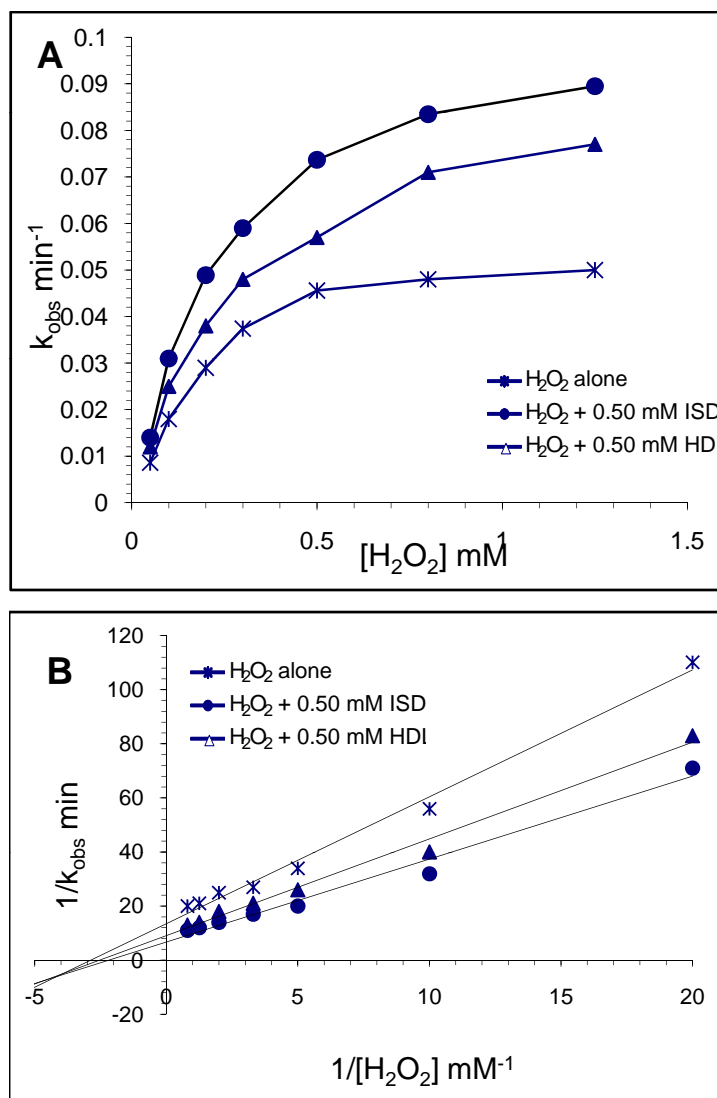


Fig. (3). Concentration dependence of H_2O_2 on MPO inactivation rate in the presence and absence of ISD and HDL. Panel A: First order inactivation rates (k_{obs}) were plotted versus the concentration of H_2O_2 . The series represents the pattern in the absence and presence of 0.5 mM ISD and HDL. Each point represents the mean of three independent determinations. Panel B: Kitz-Wilson plots for inactivation of MPO by H_2O_2 with and without 0.5 mM hydrazide. The data in panel A were transformed and plotted as double reciprocals for the determination of k_{inact} and K_I and the effect of hydrazides on those constants. MPO (20 nM) was incubated with the appropriate H_2O_2 concentration in 100 mM phosphate buffer, pH 7.4 at 20 °C in a 1.0 ml reaction volume. First order inactivation rates (k_{obs}) were determined as described in Fig (1) above and the experiments were carried out in the absence and presence of 0.5 mM hydrazides.

Consequent upon results presented in Figs. (1 and 2) that H_2O_2 alone could cause irreversible inactivation of MPO independent of the radical-generating substrates, the kinetics of H_2O_2 -dependent inactivation of MPO was investigated in the presence of the two hydrazides, ISD and HDL. Fig. (3A) shows the effects of H_2O_2 concentration on the pseudo first order inactivation rates in the presence and absence of ISD and HDL respectively. The dependence of inactivation rate constants on H_2O_2 concentration with or without the hydrazides showed a hyperbolic trend. This pattern of inactivation rate enhancement can be compared to that obtained with an activator stimulating an enzyme-catalyzed reaction [19]. It is conceivable that the mechanism of rate enhancement could be described in precise kinetics terms with defined 'activation constants' (Olorunniyi *et al.*, manuscript in preparation).

In order to obtain kinetic parameters K_I and k_{inact} for the inactivation of MPO by H_2O_2 and the effects of the hydrazides on those constants, data from Fig. (3A) were plotted as double reciprocals and shown in Fig. (3B). The resulting straight lines are similar to those obtained from the effect of an activator on an enzyme [19].

Rate enhancement by ISD and HDL involved increasing the maximum inactivation rate constant (k_{inact}) due to H_2O_2 from 0.074 min⁻¹ to 0.149 and 0.110 min⁻¹ respectively (Table 1). On the affinity of the enzyme for hydrogen peroxide as an inactivating (suicide) substrate, the inactivator dissociation constant, K_I was increased from 0.347 mM to 0.457 and 0.393 mM by ISD and HDL respectively. These results indicate that the hydrazides actually decreased the

Table 1. Rate Constants for H₂O₂-Dependent Inactivation of MPO in the Absence and Presence of ISD and HDL. Values are Mean ± Standard Deviation from Four Experiments

	H ₂ O ₂	H ₂ O ₂ + ISD	H ₂ O ₂ + HDL
k_{inact} (min ⁻¹)	0.074 ± 0.009	0.149 ± 0.012	0.110 ± 0.012
K_i (mM)	0.347 ± 0.035	0.457 ± 0.055	0.393 ± 0.018
k_{inact}/K_i (min ⁻¹ mM ⁻¹)	0.213 ± 0.029	0.326 ± 0.042	0.280 ± 0.052

affinity of MPO for H₂O₂ under inactivation conditions. Hence, it appears that while the hydrazides enhanced the maximum inactivation rate, they did not stimulate the binding and conversion of H₂O₂ to the reactive inactivating species by MPO at its active site. The hydrazides also stimulated the inactivation rate efficiency or ‘inactivation specificity constant’ (k_{inact}/K_i) due to H₂O₂, the value increasing from 0.213 min⁻¹ mM⁻¹ to 0.326 and 0.280 min⁻¹ mM⁻¹ by ISD and HDL respectively.

Table 2. Bimolecular Inactivation Rate Enhancement Factors for ISD and HDL Stimulation of H₂O₂-Dependent Inactivation of MPO. Values are Mean ± Standard Deviation from Four Experiments

	ISD	HDL
k_{enh} (min ⁻¹ mM ⁻¹)	0.184 ± 0.032	0.091 ± 0.008
k_{enh}^* (min ⁻¹ mM ⁻¹)	0.159 ± 0.047	0.088 ± 0.014

DISCUSSION

Hydrazides Enhancement of H₂O₂-Dependent Inactivation of Myeloperoxidase

The results of our investigation show that presence of a donor substrate is not a requirement for the inactivation of MPO. Kinetics analysis of the H₂O₂-dependent loss of activity is consistent with an irreversible inactivation mechanism in contrast to “reversible inactivation” involving trapping of the enzyme as compound III [18]. In a series of studies on suicide inactivation of MPO by benzoic acid hydrazides [7, 9, 20], it was suggested that a donor substrate-dependent mechanism is the main pathway for irreversible inactivation of MPO. The evidence here suggests that MPO is inactivated in a mechanism-based manner by H₂O₂, to a finite extent. The presence of the hydrazide, ISD or HDL, greatly accelerated the inactivation process. This finding is consistent with earlier studies on MPO inactivation [8, 10, 20] where the rates and extents of inactivation were markedly stimulated by the presence of a reductant substrate capable of generating inactivating intermediates or products.

We used steady state kinetics analysis to study the enhancement of H₂O₂-mediated inactivation of MPO by hydrazide donor substrates. The concentration-dependence of H₂O₂ on MPO inactivation has the same pattern in the absence and presence of the two hydrazides studied. Such pattern indicates at least a two-step mechanism for the inactivation process: an initial binding step followed by

a step involving or leading to the inactivation. The reaction of native MPO with H₂O₂ to form compound I and the subsequent formation of the inactivating species *via* other redox intermediates of the enzyme may account for this kinetics. The observation that the presence of ISD and HDL did not change the kinetics pattern of inactivation of MPO by H₂O₂ but enhanced the rate and extent of the process suggests that inactivation by H₂O₂ alone and in combination with the hydrazides might be *via* a common mechanism. It could be that the hydrazides enhanced the generation of the same inactivating species involved in H₂O₂-dependent loss of enzyme activity. This is consistent with the inactivation kinetics constants derived from the Kitz-Wilson plots (Fig. 3B).

While the two hydrazides increased the k_{inact} of the process, both decreased the affinity of the enzyme for H₂O₂ as indicated by the increase in K_i values. This suggests some competitive effect of ISD and HDL with H₂O₂ during the inactivation of MPO. This competition may be best viewed in terms of the interaction of the reactive forms i.e. the actual inactivating species of the compounds with the susceptible intermediate of the enzyme. This is a plausible explanation since the reductants are more likely to bind after H₂O₂ has reacted with the resting enzyme to form compound I [1]. Another possible explanation is that H₂O₂ and the hydrazides could compete for reaction with either compound I or compound II. These reactions might account for the increase in K_i for H₂O₂ due to the hydrazides. Goodwin *et al.* [11] reported that ISD enhanced the H₂O₂-dependent inactivation of horseradish peroxidase. The report suggested that inactivation rate enhancement by ISD may involve a mechanism similar to that due to H₂O₂ alone. ISD caused the accumulation of the enzyme as P670. This intermediate has been associated with the H₂O₂-dependent irreversible inactivation of horseradish peroxidase in the absence of reducing substrate [21].

The apparently low susceptibility of MPO to H₂O₂ inactivation reported here and elsewhere [8, 10, 20] might be due to its unusual or atypical stoichiometry in its reaction with the oxidant. MPO compound I formation requires a high concentration of H₂O₂. This behaviour has been attributed to the reversibility of compound I reaction [22] and to the ability of compound I to oxidize H₂O₂ to compound II [22, 23]. This requirement for a relatively high H₂O₂ concentration may also be linked to the pseudo-catalase activity of the enzyme [24, 25]. Steady conversion of H₂O₂ to water and molecular oxygen may further protect the enzyme from inactivation under conditions in which Ferrous MPO is

formed. Reaction of Ferrous MPO with oxygen will direct the enzyme along the turnover rather than the inactivation route [20]. The unusual covalent attachment of haem to the protein moiety of the enzyme may confer some resistance to H₂O₂-dependent inactivation.

Inactivation Rate Enhancement Factors

We have used two parameters to describe rate enhancement of MPO suicide inactivation by donor substrates: Inactivation Rate Enhancement Factor (k_{enh}) and the Specific Inactivation Rate Enhancement Factor (k^*_{enh}). The former, k_{enh} , was obtained from the slope of the plot of k_{obs} versus ISD or HDL concentrations (Figs. 1B and 2B). While k_{enh} gives an indication of the extent to which the hydrazides stimulate the inactivation rates due to H₂O₂, it confounds the inactivation enhancement caused by the hydrazide with inactivation due to H₂O₂ alone. We attempt to resolve the two effects by describing a new kinetics parameter, the specific inactivation rate enhancement factor (k^*_{enh}). This involves obtaining $k_{\text{hyd}}-k_{\text{per}}$ by subtracting the first order inactivation rates in the absence of the respective hydrazides from those determined at different hydrazides concentrations. The plots shown in Figs. (1C and 2C) are used in estimating the specific inactivation enhancement rates, by plotting $k_{\text{hyd}}-k_{\text{per}}$ versus the concentrations of ISD and HDL respectively. From the slopes of the plots shown in Figs. (1C and 2C), k^*_{enh} values were obtained. A similar approach has been used in analyzing the effects of a non-essential activator on an enzyme system in which there is finite activity in the absence of the activator [19]. Such procedure could be useful in analyzing inactivation rate enhancement by donor substrates since H₂O₂ alone caused some inactivation of the enzyme.

Other hydroperoxides that serve as oxidant for peroxidases also cause inactivation at high concentrations [1]. Hence, a kinetics parameter that defines the specific contribution of reductants to hydroperoxide-initiated inactivation of peroxidases is of significant mechanistic interest. The specific inactivation rate enhancement factor (k^*_{enh}) defined in this study could be a simple parameter for describing the rate enhancing effect of a reductant such as the two hydrazides considered here. It would be interesting to see if a given inactivating donor substrate will have comparable k^*_{enh} values when its effect is studied with different peroxides. We have preliminary evidence that analysis of protection of horseradish peroxidase from inactivation by efficient donor substrates using k_{enh} and k^*_{enh} provide interesting mechanistic insight into inactivation protection mechanisms (Iniaghe *et al.*, unpublished observation). Peroxidases have diverse industrial applications [5]. However, their inactivation by H₂O₂ even at catalytic concentrations and by the oxidation products of the donor substrates limit their practical applications. Studies on how the enzymes are inactivated are essential to understanding how to optimize their industrial uses. It is equally important to investigate and describe strategies to minimize inactivation during catalysis [5, 26]. The new kinetics parameters defined here would be relevant in the study of mechanism-based inactivation of peroxidases and related systems [27]. The kinetics constants for the reducing

substrate may provide valuable information in structure activity studies of inactivating substrates.

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